

1. Consider the following fictitious elements. They have been assigned the symbol "A" and identified using atomic notation:



- a) Which are isotopes of the same element?

- b) How many protons do these isotopes have?

- c) How many neutrons do these isotopes have?

2. Carbon atoms exist as three different isotopes: C-12, C-13, and C-14. However, the mass shown for Carbon on the periodic table is 12.01.

- a) What is the unit used for these masses? i.e. 12...what?

Ans: _____

- b) Which number/s given above (12, 12.01, 13, and 14) is/are mass number/s and which is/are atomic mass number/s?

mass number/s: _____

atomic mass number/s: _____

2. In a sample of pure uranium,
the mass number of 0.005% of the atoms is 234u
the mass number of 0.7% of the atoms is 235u
the mass number of 99.3% of the atoms is 238u

Calculate the average atomic mass of Uranium.

Answer: _____