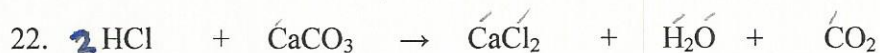
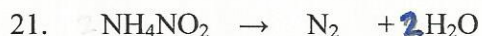
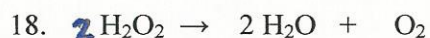
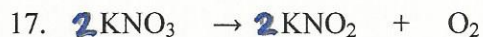
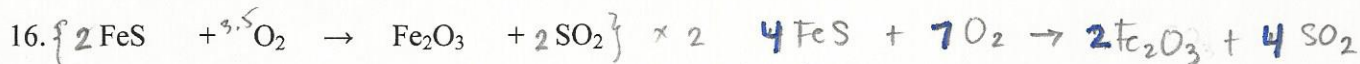


Julia Ann Cabuang  
June 9, 2015  
Tuesday

SCT 406

### Balancing Equations Practice #1

- $2 \text{SO}_2 + \text{O}_2 \rightarrow 2 \text{SO}_3$
- $4 \text{Fe} + 3 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3$
- $\text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
- $\text{TiCl}_4 + \text{O}_2 \rightarrow \text{TiO}_2 + 2 \text{Cl}_2$
- $\text{Ba}(\text{OH})_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2 \text{NaOH}$
- $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
- $\text{C}_6\text{H}_{12}\text{O}_6 + 6 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$
- $\text{Pb}(\text{NO}_3)_2 + 2 \text{KI} \rightarrow \text{PbI}_2 + 2 \text{KNO}_3$
- $\text{Cl}_2 + 2 \text{Na} \rightarrow 2 \text{NaCl}$
- $\{ 3 \text{NaOH} \rightarrow 3 \text{Na} + 1.5 \text{O}_2 + 1.5 \text{H}_2\text{O} \} \times 2 \quad 6 \text{NaOH} \rightarrow 6 \text{Na} + 3 \text{O}_2 + 3 \text{H}_2\text{O}$
- $2 \text{NH}_3 + 3 \text{Cl}_2 \rightarrow 6 \text{HCl} + \text{N}_2$
- $\{ \text{N}_2 + 2.5 \text{O}_2 \rightarrow \text{N}_2\text{O}_5 \} \times 2 \quad 2 \text{N}_2 + 5 \text{O}_2 \rightarrow 2 \text{N}_2\text{O}_5$
- $\text{N}_2 + 3 \text{H}_2 \rightarrow 2 \text{NH}_3$
- $\{ 2 \text{P} + 2.5 \text{O}_2 \rightarrow \text{P}_2\text{O}_5 \} \times 2 \quad 4 \text{P} + 5 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_5$

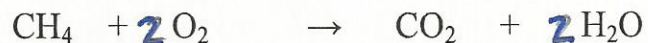


25. The neutralization of hydrochloric acid (HCl) by calcium carbonate (CaCO<sub>3</sub>) produces calcium chloride (CaCl<sub>2</sub>), carbon dioxide (CO<sub>2</sub>) and water (H<sub>2</sub>O).

Write the balanced equation for this neutralization reaction.



26. The combustion of methane, CH<sub>4</sub> mixed with O<sub>2</sub> produces carbon dioxide, CO<sub>2</sub>, and water, H<sub>2</sub>O. The unbalanced chemical equation for this reaction is as follows:



Balance the chemical equation for this combustion reaction.