

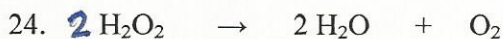
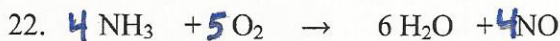
Jule Ann Cabuang
Tuesday
June 09, 2015

SCT 406

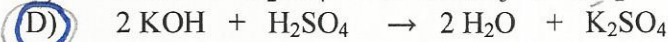
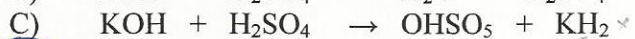
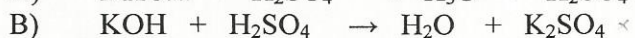
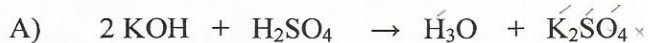
Balancing Equations Practice #2

- $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
- $2\text{KClO}_3 \rightarrow 3\text{O}_2 + 2\text{KCl}$
- $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$
- $3\text{Hf} + 2\text{N}_2 \rightarrow \text{Hf}_3\text{N}_4$
- $5\text{HNO}_3 + \text{P} \rightarrow \text{H}_3\text{PO}_4 + 5\text{NO}_2 + \text{H}_2\text{O}$
- $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$
- $\text{C}_2\text{H}_5\text{OH} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$
- $2\text{N}_2 + 5\text{O}_2 \rightarrow 2\text{N}_2\text{O}_5$
- $\text{C}_2\text{H}_2 + 2\text{H}_2 \rightarrow \text{C}_2\text{H}_6$
- $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$
- $6\text{O}_2 + \text{C}_4\text{H}_9\text{OH} \rightarrow 4\text{CO}_2 + 5\text{H}_2\text{O}$
- $6\text{Sr} + 2\text{N}_2 \rightarrow 2\text{Sr}_3\text{N}_2$
- $2\text{Cu}_2\text{S} + 2\text{O}_3 \rightarrow 2\text{Cu}_2\text{O} + 2\text{SO}_2$
- $\text{Mg}_3\text{N}_2 + 3\text{H}_2\text{O} \rightarrow 3\text{MgO} + 2\text{NH}_3$
- $\text{Cr}(\text{OH})_3 + \text{NaOH} \rightarrow \text{NaCrO}_2 + 2\text{H}_2\text{O}$
- $3\text{FeCl}_2 + 2\text{Na}_3\text{PO}_4 \rightarrow \text{Fe}_3(\text{PO}_4)_2 + 6\text{NaCl}$
- $6\text{HCl} + 2\text{Al} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$
- $2\text{H}_2\text{SO}_4 + \text{C} \rightarrow \text{CO}_2 + 2\text{H}_2\text{O} + 2\text{SO}_2$
- $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$
- $2\text{NaI} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{PbI}_2 + 2\text{NaNO}_3$

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29. Choose the balanced equation



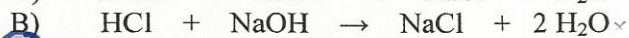
a) $\text{S} = 1$
 $\text{K} = 2$
 $\text{O} = 6$
 $\text{H} = 4$

$\text{O} = 5$
 $\text{H} = 3$

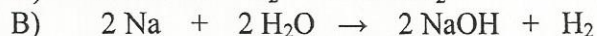
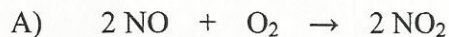
b) $\text{K} = 2$
 $\text{O} = 6$
 $\text{H} = 4$
 $\text{S} = 1$

$\text{K} = 2$
 $\text{O} = 6$
 $\text{H} = 4$
 $\text{S} = 1$

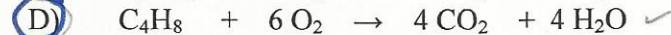
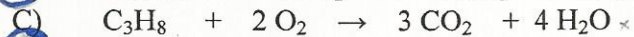
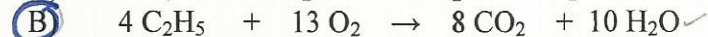
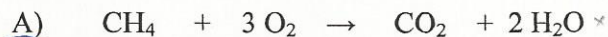
30. Choose the balanced equation



31. Which equation is **not** balanced?



32. Which **equations** are balanced?



33. The complete reaction of 8 g of methane (CH_4) with 71 g of chlorine gas (Cl_2) produces 73 g of hydrochloric acid (HCl) and a certain amount of carbon (C).

The balanced equation for this reaction is as follows:

