

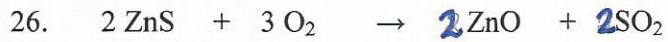
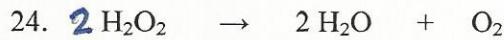
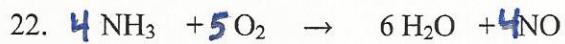
Julie Ann Cabuay
Tuesday
June 09, 2015

SCT 406

Balancing Equations Practice #2

1. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
2. $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$
3. $2\text{KClO}_3 \rightarrow 3\text{O}_2 + 2\text{KCl}$
4. $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$
5. $3\text{Hf} + 2\text{N}_2 \rightarrow \text{Hf}_3\text{N}_4$
6. $5\text{HNO}_3 + \text{P} \rightarrow \text{H}_3\text{PO}_4 + 5\text{NO}_2 + \text{H}_2\text{O}$
7. $\text{P}_4\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$
8. $\text{C}_2\text{H}_5\text{OH} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 3\text{H}_2\text{O}$
9. $2\text{N}_2 + 5\text{O}_2 \rightarrow 2\text{N}_2\text{O}_5$
10. $\text{C}_2\text{H}_2 + 2\text{H}_2 \rightarrow \text{C}_2\text{H}_6$
11. $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$
12. $6\text{O}_2 + \text{C}_4\text{H}_9\text{OH} \rightarrow 4\text{CO}_2 + 5\text{H}_2\text{O}$
13. $6\text{Sr} + 2\text{N}_2 \rightarrow 2\text{Sr}_3\text{N}_2$
14. $2\text{Cu}_2\text{S} + 2\text{O}_3 \rightarrow 2\text{Cu}_2\text{O} + 2\text{SO}_2$
15. $\text{Mg}_3\text{N}_2 + 3\text{H}_2\text{O} \rightarrow 3\text{MgO} + 2\text{NH}_3$
16. $\text{Cr}(\text{OH})_3 + \text{NaOH} \rightarrow \text{NaCrO}_2 + 2\text{H}_2\text{O}$
17. $3\text{FeCl}_2 + 2\text{Na}_3\text{PO}_4 \rightarrow \text{Fe}_3(\text{PO}_4)_2 + 6\text{NaCl}$
18. $6\text{HCl} + 2\text{Al} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$
19. $2\text{H}_2\text{SO}_4 + \text{C} \rightarrow \text{CO}_2 + 2\text{H}_2\text{O} + 2\text{SO}_2$
20. $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$
21. $2\text{NaI} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{PbI}_2 + 2\text{NaNO}_3$

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29. Choose the balanced equation

- A) $2 \text{KOH} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_3\text{O} + \text{K}_2\text{SO}_4 \times$
 B) $\text{KOH} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{K}_2\text{SO}_4 \times$
 C) $\text{KOH} + \text{H}_2\text{SO}_4 \rightarrow \text{OHSO}_5 + \text{KH}_2 \times$
 D) $2 \text{KOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{K}_2\text{SO}_4$

a)	$\text{S} = 1$	$\text{O} = 5$
	$\text{K} = 2$	$\text{H} = 3$
	$\text{O} = 6$	
		$\text{H} = 3$
		$\text{S} = 1$

30. Choose the balanced equation

- A) $2 \text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O} \times$
 B) $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + 2 \text{H}_2\text{O} \times$
 C) $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O} \checkmark$
 D) $\text{HCl} + 2 \text{NaOH} \rightarrow 2 \text{NaCl} + 2 \text{H}_2\text{O} \times$

31. Which equation is not balanced?

- A) $2 \text{NO} + \text{O}_2 \rightarrow 2 \text{NO}_2$
 B) $2 \text{Na} + 2 \text{H}_2\text{O} \rightarrow 2 \text{NaOH} + \text{H}_2$
 C) $\text{H}_3\text{PO}_4 + 3 \text{KOH} \rightarrow \text{K}_3\text{PO}_4 + 3 \text{H}_2\text{O}$
 D) $3 \text{HBr} + \text{Fe(OH)}_3 \rightarrow \text{FeBr}_3 + 6 \text{H}_2\text{O}$

32. Which equations are balanced?

- A) $\text{CH}_4 + 3 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O} \times$
 B) $4 \text{C}_2\text{H}_5 + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O} \checkmark$
 C) $\text{C}_3\text{H}_8 + 2 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O} \times$
 D) $\text{C}_4\text{H}_8 + 6 \text{O}_2 \rightarrow 4 \text{CO}_2 + 4 \text{H}_2\text{O} \checkmark$

33. The complete reaction of 8 g of methane (CH_4) with 71 g of chlorine gas (Cl_2) produces 73 g of hydrochloric acid (HCl) and a certain amount of carbon (C).

The balanced equation for this reaction is as follows:

