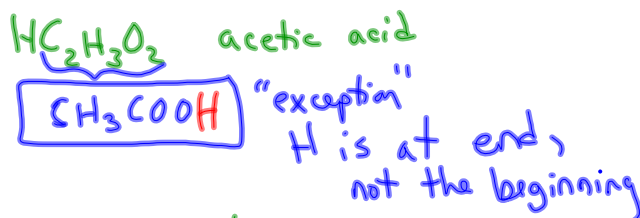
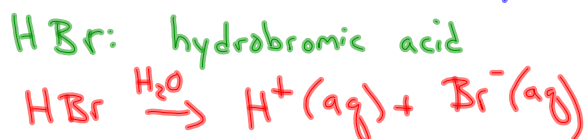
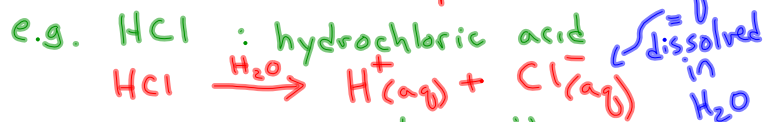


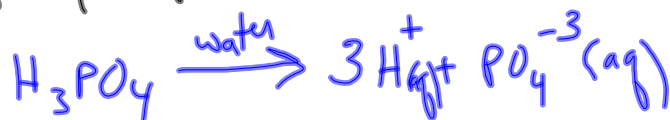
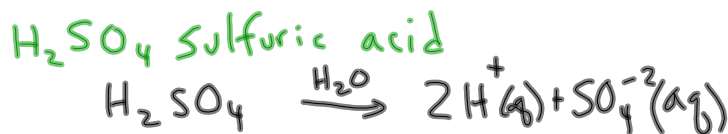
Acids, Bases, and Salts

Acid: All acids are ionic compounds that contain H and an anion ^{non-metal portion} _(- ion) "aqueous"
^{metal portion}

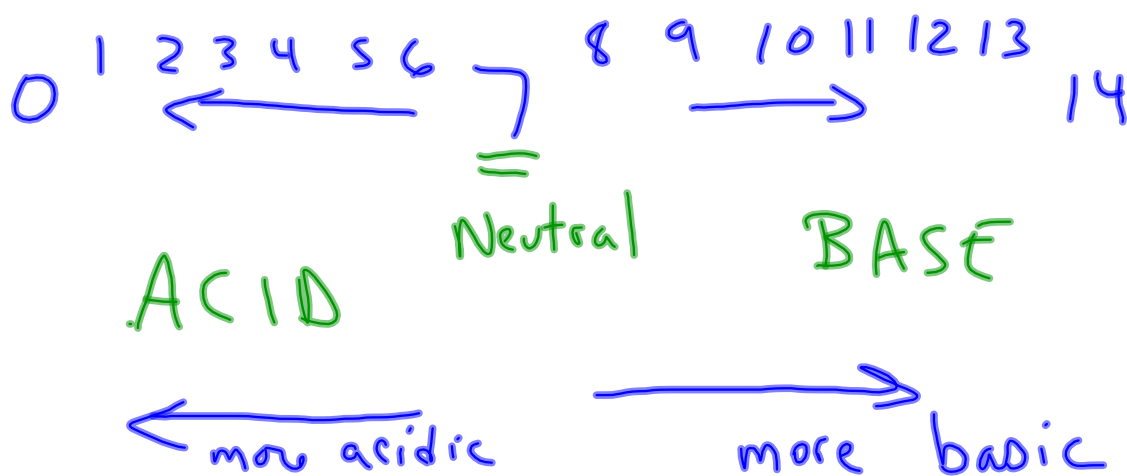


We know this is an acid because it dissociates in water to release H^+ .

[↑] "hydronium ion"



pH Scale 0 - 14



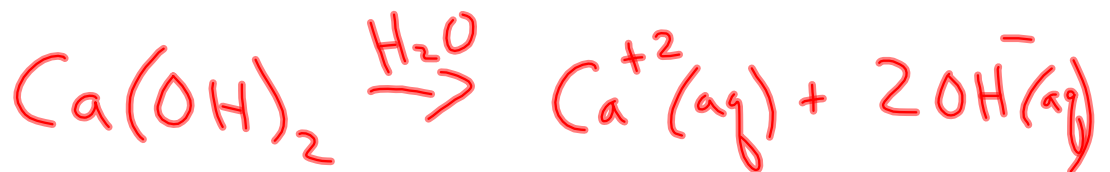
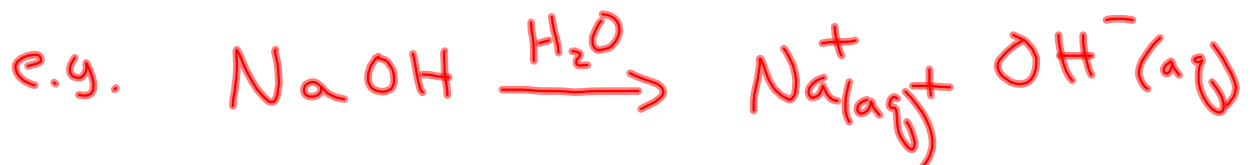
Litmus

blue litmus would turn red
red litmus would stay red

ACID = RED

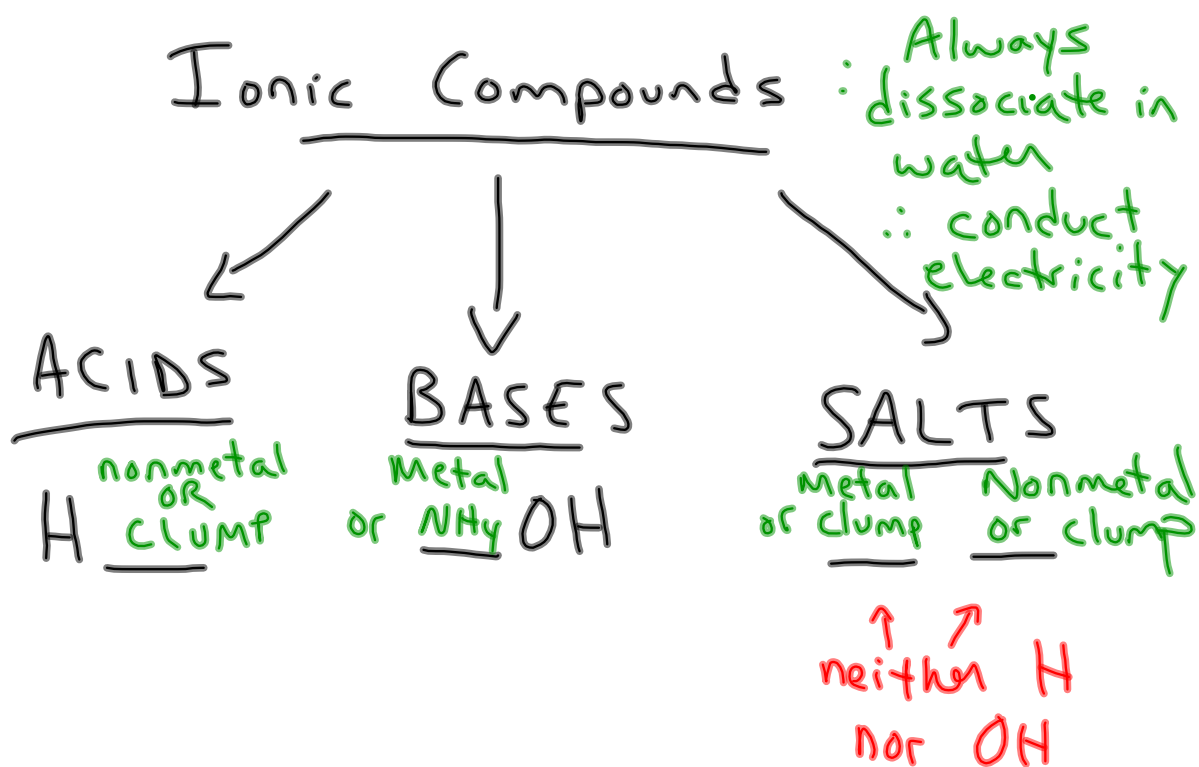
Bases

→ dissociate in water to
release OH^- (hydroxide)

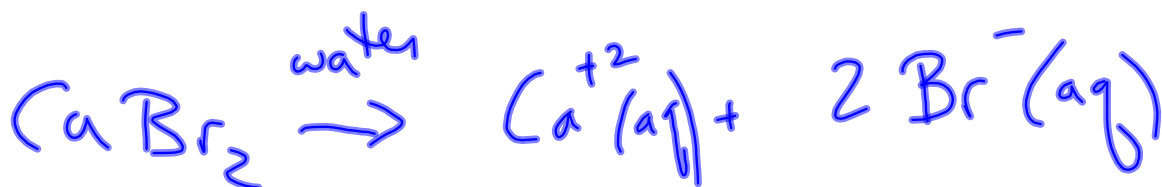
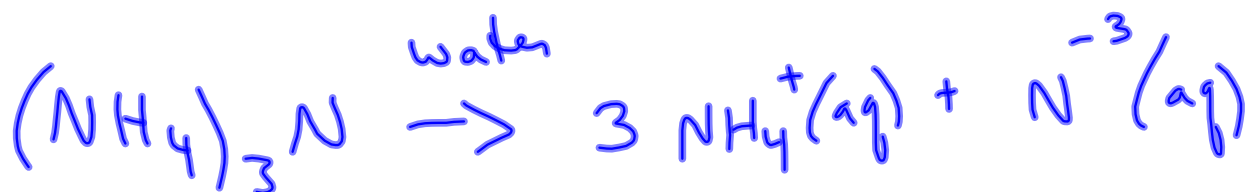
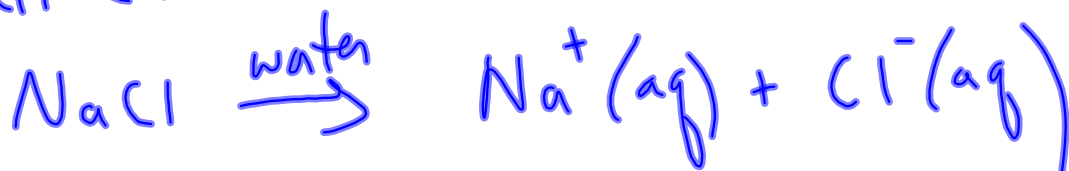


Bases turn red litmus blue;
blue litmus will stay blue.

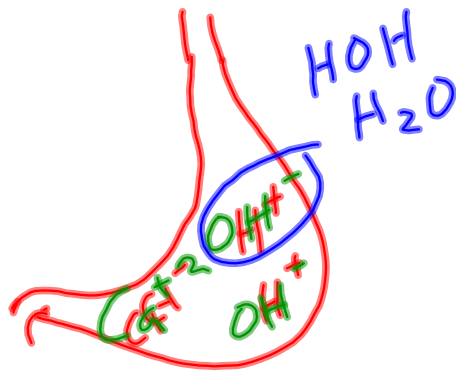
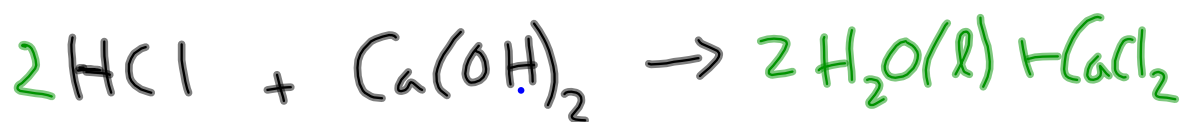
pH Bases: 7 → 14
(not incl. 7)



Salt dissociation :



Acids + Bases will Neutralize
Each Other



Reason for acid:

dissociates in water to
release H^+

Reason for base:

dissociates in water to
release OH^-

Reason for salt: dissociates in
water but doesn't release H^+ or OH^-

a) $\begin{matrix} \text{blue} \rightarrow \text{blue} \\ \text{red} \rightarrow \text{red} \end{matrix}$ neutral and weak electrolyte? **G**

b) B

e) E

c) A

f) F

d) D

g) H

2. C