PSC-4010

Using the Periodic Table to Determine Information on Elements

Examine the electron configuration diagrams that you drew for the first 20 elements. Respond to the following questions:

1. What is the relationship between the Group # of an element and the number of valence electrons that the element has?

2. What is the relationship between the period number of an element and the number of energy levels (orbits) that the element has?

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Information on Magnesium

Symbol: _____ Atomic number: _____

Atomic mass: _____ Number of protons: _____

Number of electrons: _____ Number of neutrons: _____

Group number: _____ Period number: _____

Number of energy levels: _____

Number of electrons in the outermost energy levels: _____

Information on Fluorine

Symbol: _____ Atomic number: _____

Atomic mass: _____ Number of protons: _____

Number of electrons: _____ Number of neutrons: _____

Group number: _____ Period number: _____

Number of energy levels: _____

Number of electrons in the outermost energy levels: _____

Information on Phosphorus

Symbol: _____ Atomic number: _____

Atomic mass: _____ Number of protons: _____

Number of electrons: _____ Number of neutrons: _____

Group number: _____ Period number: _____

Number of energy levels: _____

Number of electrons in the outermost energy levels: _____

Information on Sodium

Symbol: _____ Atomic number: _____

Atomic mass: _____ Number of protons: _____

Number of electrons: _____ Number of neutrons: _____

Group number: _____ Period number: _____

Number of energy levels: _____

Number of electrons in the outermost energy levels: _____